

MOGA, A., Acad.; OBRASCU, C., dr.; DOBO, dr.; TOMAS, Alexandrina, prof.,
ed., fiz.; BLENDEA, O, dr., si colectivul.

Study of medical physical therapy of hypertensive disease.
Med. int., Bucur. 4 no.8:1177-1181 Dec 56.

(HYPERTENSION, therapy
phys. ther.)

(PHYSICAL THERAPY, in various dis.
hypertension)

RUMANIA

NICOLAU, Cl., Conf. Dr., TONAS, E.; OLINESCU, R.; CHRISTEA, Al., CONSTANTINESCU, Rodica; and STROESCU, Eugenia

"Activity of 2-Methyl-1, 4-Naphthoquinone Sodium Bisulfite(Vitamin K3) on Some Redox Enzymes"

Bucharest, Revista Sanitara Militara, Vol 16, Special No., 1965; pp 389-393

Abstract: In vitro studies to pinpoint mode of radiosensitizing effect of Vitamin K3 in study with ceruloplasmin, catalase, peroxidase, d-amino-acidoxidases. Results indicate that K3 vitamin has profound effect inhibiting or potentiating the enzymatic activities depending on its concentration. This is probably the mode of action of Vitamin K3 as radiosensitizer.

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- 78 -

TOMAS, G.Ya., inzh..

Prospects for the wide use of floating factories in the national
economy of the U.S.S.R. Stroi. dor. mash. 8 no.5:19-22 My '63.

(MIRA 16:5)

(Factories—Design and construction)

(Construction industry)

ZAVALISHIN, I., inzh.; TOMAS, G., inzh.

Floating dwellings for builders. Znil. stroi. no.9:31 '64.
(MIRA 17:12)

TCMAS, G. Ya.

Reinforced concrete constructions. Mashinostroitel' no.3:11-12 Mr '61.
(MIRA 14:3)

(Reinforced concrete constructions)

BRUCKNER, L.; MOSLER, J.; TOMAS, J.; CERNY, J.; BESKA, F.

X-ray picture of the urinary bladder in gynecological carcinoma.
Cesk.rentg.14 no.6:390-395 D'60.

1. Onkologické odd. KUNZ-Ostrava v Paskově, přednosta MUDr.
B. Roffersberg.

(BLADDER radiog)

(GENITALIA FEMALE neopl)

TOMAS, Jaroslav

CO-CO₂ ionization analyzer. Jaderna energie 9 no. 12: 391
D '63.

1. Katedra analytické chemie a radiochemie, Vysoká škola
banská, Ostrava.

TOMAS, Jaroslav

Development of the apparatus for measurement of thermodynamic constants of alloys. Jaderna energie 9 no.11: 357 '63.

1. Katedra analytické chemie a radiochemie, Vysoká škola báňská, Ostrava.

TOMAS, Jaroslav, inz.

Protection of synchronous machines in case of excitation loss.
Energetika Cz 11 no.1:36-38 Ja '61.

SLAVIK, J. B., prof., RNDr.; TOMAS, J.

The 2nd Conference on Acoustics in Budapest. Slaboproudy obzor 22
no.12:778 D '61.

(Sound)

L 46899-66

ACC NR: AP6034285

SOURCE CODE: CZ/0034/66/000/005/0329/0333

AUTHOR: Tomas, Jindrich (Engineer)

20

ORG: Chair of Energetics, College of Mining, Ostrava (Katedra Energetiky, Vysoka Skola Banska)

B

TITLE: Some shortcomings in the operation of soaking pit furnaces

SOURCE: Hutnicke listy, v. 21, no. 5, 1966, 329-333

TOPIC TAGS: furnace, metallurgy

ABSTRACT:[Author's English summary modified]: Operation of refractory regenerators of the soaking pit furnaces is analyzed and experimental results are evaluated statistically. Part of the air heated in the regenerators is lost by cracks and holes in the lining; some escape to flue gases, and some to atmosphere. The amount lost changes during the course of the run, and in individual furnace runs. The losses result in a higher oxygen content of flue gases than the ratio corresponding to the control instruments settings. Combustion characteristics of flue gases, and their thermal properties are described. The generally used four burner system in soaking pit firing is evaluated. Orig. art. has: 11 figures and 6 tables. [JPRS: 36,867]

SUB CODE: 13 / SUM DATE: none

Card 1/1 fv

UDC: 621.783.224

092/0006

~~TOMAS, J.~~

TECHNOLOGY

periodicals: POZEMNI STAVBY Vol. 7, no. 2, Feb. 1959

TOMAS, J. Building of factory chimneys. p. 64.

Monthly List of East European Accession (EEAI) LC Vol. 8, no. 5
May 1959, Unclass.

TOMAS, Jaroslav, inz., TRINDL, Josef, prof. inz. DrSc.

Using the fast radiometric method for the analysis of tin
solders and for the control of coating with tin-lead alloys.
Sit listy 19 no. 6:430-432 Je '64.

1. Higher School of Mining, Ostrava.

TOMAS, Jaroslav, inz.

Earth connection of generator rotora. Energetika Cz 11 no.2195"
98 F '61.

BRUCKNER, Ladislav; BESKA, Frantisek; MOSLER, Jiri; TOMAS, Jaroslav

The kidneys and ureters in carcinoma of the rectum and sigmoid.
Roshl. chir. 40 no.12:844-850 '61.

1. Onkologicke oddeleni KUNZ Ostrava v Paskove, prednosta MUDr.
B. Roffersberg.

(RECTUM neoplasms)	(SIGMOID neoplasms)
(KIDNEY radiography)	(URETER radiography)

TOMAS, Jan. Adv. Inz.: GUTKA, A., Inz.; KUCHAR, Jan. Inz.; HANEC, Otakar, Inz.; TEINDL, Josef, prof. Inz. Inz.

Use of the radiocapture method in the study of corrosion, especially in metal coatings. Sbor. VGB Olomava 10 no.3:341-349 '64.

1. Corresponding member of the Czechoslovak Academy of Sciences (for Teindl). Submitted June 8, 1964.

TOMAS, Jaroslav, inz.

Apparatus for measuring metal tension and thermodynamic activities at high temperatures. Sbor VSB Ostrava 10 no.3: 319-327 '64.

1. Submitted June 8, 1963.

SOUKUP, Lubomir; MARTINEK, Ladislav; TOMAS, Josef, inz.

New method of manufacturing flywheel gear rims. Automobil
Cz 8 no.1:24-25 Ja '64.

1. Automobilove zavody, Mlada Boleslav.

LIST AND THE COUNTRY																									
PROCESSES AND PROPERTIES INDEX																									
<p>Recent find of sulfur and salammuniac at Kolcany in Moravia. Tomáš Krulík. <i>Věda Pěstová, Praha</i> 23, 61 (1944); <i>Minerology</i> 18:17-19; 203 (1946). All 3 modifications of S, salammuniac, and numerous sulfates formed by the oxidation of marcasite are reported from this lignite mine. Michael Fleischer</p>																									
<p>ASB-SEA METALLURGICAL LITERATURE CLASSIFICATION</p>																									
<p>STONY DIVISION</p>																									
<p>STONY DIVISION</p>																									

TAYLER, D.K.; TOMAS, L.F.; SHERIDAN, D.

Comments on the article "Rotational spectrum of the cyanamide
molecule. Opt. i spektr. 12 no.3:452 Mr '62. (MIRA 15:3)
(Cyanamide--Spectra) (Molecular rotation)

MITACHI, N., coresp.; IOSIF, B., coresp.; VALKAY, Geza, coresp.; TOMAS,
Liviu, coresp.

In short. Constr Buc 17 no.793:4 20 Mr '65.

VALKAI, Geza, coresp.; TOMAS, Liviu, coresp.

Balance of a trimester rich in achievements. Constr Buc 17
no.798:1 24 Ap '65.

BERGER, Vladimir; TOMAS, Michal

Urea adhesives with a low content of free formaldehyde. Drevo
19 no.6:211-214 Je '64.

1. State Research Institute of Wood, Bratislava.

TOMAS, Michal; JURIK, Ladislav

Urea-formaldehyde resins with low content of free formaldehyde.
Drevarsky vysekum no.3:171-178 '63.

1. Statny drevarsky vyskumny ustav, Bratislava.

TOMAS, O.; RUSHCHINSKIY, L.

Consolidate the planning departments, improve the planning. Mor.
flat. 24 no.8:11 Ag '64. (MIRA 18:9)

1. Nachal'nik Odesskogo porta (for Tomas). 2. Nachal'nik planovogo
otdela Odesskogo porta (for Rushchinskiy).

TOMAS, O.

Industrial potentialities within harbors to serve the seven-year plan. Mor. flot 23 no.8:12-13 Ag '63. (MIRA 16:11)

1. Nachal'nik Odesskogo porta.

TOMAS, O.

Car exchange area is a direct variant of growth potentiality.

Mdr. flot 23 no.3:8-10 Mr '63.

(MIRA 16:3)

1. Nachal'nik Odesskogo porta.

(Odessa—Cargo handling)

(Railroads—Freight cars)

TOMAS, Petar, dr

The neutron reactions at 14 Mev.; abstract of a doctoral dissertation. Glas mat fiz Hrv 16 no.3/4:325-326 '61.

PAIC, V.; PAIC, M.; PRELEC, K.; CERINEO, M.; ILAKOVIC, K.; SLAUS, I.; TOMAS, P;
VALKOVIC, V.; LJOLJE, K.; SIPS, V.

Review of periodicals; physics. Bul sc Youg 9 no.4/5:126 Ag-0
'64.

1. Ruder Boskovic Institute, Zagreb.

ANTOLKOVIC, B. (Zagreb); PAIC, M. (Zagreb); PRELEC, K. (Zagreb);
TOMAS, P. (Zagreb); TURK, M. (Zagreb); WINTERHALTER, D. (Zagreb)

The absolute and relative measurements of neutron fluxes obtained
from the neutron generator of the Institute "Ruder Boskovic."
Ves mat fiz Srb no.12:97-101 '60.

TOMAS, Petar (Zagreb)

Production of thin films by thermal evaporation. Gl mat fiz Hrv 15
no.2:119-134 '60. (EEAI 10:9)

1. Institute "Ruder Boskovic", Zagreb.

(Evaporation) (Thin films) (Metallic films)

YUGOSLAVIA/Nuclear Physics - Installations and Instruments.
Methods of Measurement and Research

C-2

Abs Jour : Ref Zhur - Fizika, No 2, 1959, No 2599

Author : Paic M., Prelec K., Tomas P., Varicak M., Vosicki B.

Inst : -

Title : Cockroft and Walton Accelerator for 200 kb Used to Generate
Neutrons.

Orig Pub : Glasnik mat.-fiz. i astron., 1957, 12, No 4, 269-289

Abstract : No abstract

Card : 1/1

STERNBERG, Z.; TOMAS, P.

Excitation of helium atoms by the impact of deuterons and
rotons. Bul sc Youg 7 no.1/2:19 F-Apr '62.

1. Institut "Ruder Boskovic," Zagreb.

STIPCIC, N. (Zagreb); PAIC, M. (Zagreb); TOMAS, P. (Zagreb)

The ion optical system of a 200 kV Cockroft-Walton accelerator.
Glas mat fiz Hrv 17 no.1/2:107-112 '62.[publ. '63].

1. Institute "Ruder Boskovic", Zagreb.

TOMAS, V. MUDr.

Work of the medical statistician in factory institutes of
national health. Cesk. zdrav. 12 no.4: 201-201 Ap'64

1. Zavodni ustav narodniho zdravi Novahut Klementa Gottwalda,
Ostrava -Kuncice.

CZECHOSLOVAKIA

TOMAS, V., Promoted Physician

ZUNZ-NHKG (ZUNZ-NHKG), Ostrava-Kuncice

Prague, Prakticky lekar, No 4, 1963, pp 153-156

"Twins of Czech History."

TOMAS, V.

From the history of epidemics. I. Reflections on the origins of epidemics in the 13th to the 18th centuries. Cas. lek. cesk. 103 no.32:887-888 Ag 7 '64.

1. Dilensky obvod 12 Zavodniho ustavu narodniho zdravi Novahut Klementa Gottwalda, Ostrava-Kuncice (reditel MUDr. Z. Vich).

TOMAS, V.

On the prevention of epidemics. II. Treatment of infectious diseases
in the Middle Ages. Cas. lek. cesk. 103 no.33:920-922 14 Ag '64.

1. Dilensky obvod 12 Zavodni ustav narodniho zdravi Novahut Klementa
Gottwalda, Ostrava-Kuncice (reditel MUDr. Z. Vich).

TOMASH, Vargeliy [Tomas, V.] (Vengriya)

New developments in the oldest equipment. Tekh. mol. 28 no. 3:13
'60. (MIRA 14:4)

(Hungary—Grain milling machinery)

TOMAS, V., Poruba IV, 1372

Tomas Jordan (1539-1586) in the eyes of his contemporaries, according to a commemorative medal from the year 1570. Cas. lek. Cesk. 104 no.46:1278-1279 19 N '65.

TOMAS, Vladimir

CZECHOSLOVAKIA

Graduate physician

Not given (address: Ostrava I, Privozska 23)

Prague, Prakticky Lekar, No 21, Nov 62, p 924

"Contribution to Dr. F. Potuzil's article 'Indemnification for Psychic Accidents at Work and Other Impairment of Health' (Prakticky Lekar, No 19, 1961)

TOMAS, Vladimir

SURNAME, Given Names

Country: Czechoslovakia

(2)

Academic Degrees: Graduated Physician (Promovaný lékař)

Affiliation: Kraj Institute of Public Health (Krajský ústav národního zdraví),
Ostrava V; Director: J. ČERNÝ, MD.

Source: Prague, Praktický Lékař, Vol 41, No 9, 1961, pp 423-426.

Data: "Development of Hospital Care in the Ostrava Area."

670 981643

TOMAS, Vladimir
SURNAME, Given Names

(2)

Country: Czechoslovakia

Academic Degrees: Graduate Physician (promovaný lékař)
Kraj Institute of Public Health (Krajský ústav národního
Affiliation: zdraví), Ostrava; Chief (Prednosta) of the Internal Department (interní oddělení): Dr J Cerný

Source: Prague, Praktický Lékař, Vol 41, No 17, 5 September 1961, pp 796-800

Data: "Incompetents and Quacks in the First Half of the Nineteenth Century at Ostrava."

134

070 981643

ZHUKOV, A.A., kand.tekhn.nauk; SHALASHOV, V.A., inzh.; TOMAS, V.K., inzh.

The structure of cementite. Lit. proizv. no.7:46 JI '65.
(MIRA 18:8)

SHALASHOV, V.A.; Primali uchastiye: ZHUKOV, A.A.; TOMAS, V.K.

Effect of nuclear radiation on the thermodynamics of metal alloys.
Fiz. met. i metalloved. 16 no.2:278-284 Ag '63. (MIRA 16:8)

1. Vsesoyuznyy nauchno-issledovatel'skiy institut tekstil'nogo i
legkogo mashinostroyeniya.

(Metals, Effect of radiation on)
(Alloys—Thermodynamic properties)

TOMAS, V.

On the history of epidemics. III. Protection against infection
from the 13th to the 18th centuries. Cas. lek. cesk. 103
no.37:1034-1036 11 S '64.

1. Dilensky obvod 12, Zavodni ustav narodniho zdravi Novahut
Klementa Gottwalda, Ostrava-Kuncice (reditel dr. Z. Vich).

TOMAS, V. (Prom. Dr.)

CZECHOSLOVAKIA

TOMAS, V., Prom. Dr.

ZUNZ-MHKG, Ostrava-Kuncice

Prague, Prakticky lekar, No 13-14, 1963, pp 552-553

"The First Czecho Women Physicians."

SHALASHOV, V.A.; Prinimali uchastnye: BREGER, A.Kh.; ZHUKOV, A.A.; GOL'DIN, V.A.; TOMAS, V.K.

Effect of irradiation on the structure and tendency to thermal decomposition of chromium cementite. Zhur.fiz.khim. 38 no.11: 2735-2737 N '64. (MIRA 18:2)

TOMAS, V.K.

Determination of iron carbide from interplanar spacings.
Zav.lab. 31 no.4:453-455 '65.

(MIRA 18:12)

1. Vsesoyuznyy nauchno-issledovatel'skiy institut tekstil'nogo
i legkogo mashinostroyeniya.

TOMAS, Y. [Thomas, J.]

Accurate method for determining the active concentration of
natural radionuclides. Atom. energ. 12 no.5:431-433 My '62.
(MIRA 15:5)

1. Institut gigiyeny truda i profzabolevaniy, Praga.
(Radioactivation analysis) (Aerosols)

TOMAS, Zdenek, inz.

Automation of ship construction in Czechoslovakia. Doprava
6 no.5:446-450 '64.

TOMAS, Z., inz.

Ten years of the Section for Water Transport of the Transport
Research Institute. Vodní hosp 13 no.1:37 '63.

TOMAS, Zdenek, inz.; KRYSL, Frantisek, inz.

The central harbor in Usti nad Labem. Doprava no.3:88-91
'62.

CZECHOSLOVAKIA/Chemical Technology - Chemical Products and H.
Their Application. Synthetic and Natural Medicinal
Substances. Galenicals and Medicinal Forms.

Abs Jour : Ref Zhur - Khimiya, No 10, 1959, 35992
Author : Pivoda, A., Tomsch, E.
Inst : -
Title : The Application of Isotopes in the Pharmaceutical Indus-
try.
Orig Pub : Farmacia (Ceskosl.), 1958, 27, No 8, 235-240.
Abstract : The possibility of the application of isotopes for steri-
lization of drugs in ampules and for their analysis is
indicated. -- I. Matveyeva.

Card 1/1

CZECHOSLOVAKIA/Analytical Chemistry - Analysis of Inorganic Substances.

E.

Abs Jour : Ref Zhur - Khimiya, No 9, 1958, 28468

and an aliquot portion is titrated potentiometrically with 0.02 N KMnO_4 . When the first jump in potential is reached, 10 ml of 0.5 N KCN or 10 ml acetone are added to the above solution and the titration with 0.02 N KMnO_4 is continued until the second jump in potential is reached. The addition of CH_3COOH to the solution to be titrated is intended to prevent the volatilization of the I_2 . The authors reject the possibility of the formation of I^+ during the titration and are of the opinion the I_2 reacts with KCN to form ICN and I^- ; when acetone is used, $\text{CH}_3\text{COCCH}_2\text{I}$, I^- , and H^+ are assumed to be formed. 1 gm-equiv of KMnO_4 is equivalent to 2 gm-equiv I^- . In the iodometric method the solution to be analyzed is mixed with 25 ml 8 N H_2SO_4 and 50 ml glacial CH_3COOH and the resulting solution is diluted to 100 ml; an aliquot portion is then titrated with a 0.05 N solution

Card 2/3

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1st AND 2ND SERIES										PROCESS AND PROPERTIES INDEX										3RD AND 4TH SERIES									
<div style="position: absolute; top: 10px; left: 10px; font-size: 2em; font-family: cursive;">SV</div> <div style="position: absolute; top: 10px; right: 10px; font-size: 2em; font-family: cursive;">H53A</div> <div style="position: absolute; top: 300px; left: 250px;"> <p>2047. Lenard Phosphors. R. Tarnaschok. Acta Physica Polonica, 8. pp. 393-406; Disc. 406-408, 1954. In German.—A review of the present position with regard to the phosphorescent spectra emitted by Lenard-phosphors. Specially dealt with are: the relationship between the spectra observed and the atomic term-scheme of the metallic impurity, rare-earth phosphors, and the effect of the lattice-vibrations of the material on spectra emitted.</p> <p style="text-align: right;">A. C. M.</p> </div>																													
ASD-51A METALLURGICAL LITERATURE CLASSIFICATION																													
18000 SYNOPTIC										18000 ANALYTIC										18000 INDEX									
18000 ANALYTIC										18000 INDEX										18000 INDEX									

TOMASCHEK, R.

Practice and problems of earth tide measurements. Geofiz kozl
9 no.1/2:69 '60.

TOMASCHEK, Zoltan, a muszaki tudományok kandidátusa; MAKO, Zoltan; MAGYAR, Laszlo; VAMBERI, Lorinc; KONCZ, Istvan

Properties of the titanium getter and its use in electronic tubes of great specific loading; also, remarks by Z.Mako and others. Muszaki kozl MTA 26 no.1/4:219-220 '60. (EEAI 9:10)

1. Híradástechnikai Kutató Intézet (for Tomaschek)
(Electron tubes) (Titanium)

10MACHÉK, Z.

10. The following properties of compact titanium in vacuum system: Z. Machéček. Thesis for the degree of candidate of sciences. 138 p., 45 figs., 3 tabs.

The following properties of compact titanium were investigated in an apparatus made entirely of glass, without leaks or ground parts. The apparatus consisted of a Pirani vacuum meter for the 10^{-1} to 10^{-2} mm mercury range and an ionization vacuum meter for the range of 10^{-2} to 10^{-3} mm mercury, an electrically heated tube containing the piece of titanium to be investigated and several ampoules containing the pure gases used in the tests and provided with an impact seal. While being heated to 420–450°C the apparatus was evacuated to 10^{-3} mm mercury, sealed and the measurements made in this condition. The experiments were conducted at 400–800°C and at 1000°C. After breaking an ampoule the pressure drop and the time required for attaining a state of equilibrium were measured. When breaking several ampoules the amount of gas added and the time necessary to reach equilibrium were totaled. The graphic representation of the results yields a parabola from the outset for the Ti–N system, for the Ti–O system — until the absorption of about 2.5% by weight of oxygen (solid solution) — a linear function and later, after the formation of an oxide, a parabolic function. The amount of oxygen and nitrogen absorbed was determined and compared with the

results obtained under similar circumstances for compact zirconium; the volume of gases absorbed during a comparable time was considerably greater for titanium. Suction rates were also computed from the above data showing a higher rate for titanium especially after a certain volume of gases had already been absorbed by the metals. The author determined the parabolic rate constants and from these the energies of activation for the temperature ranges 1000 to 800°C and 800 to 400°C; they show an acceptable agreement with published data. The change in the specific electrical resistance was determined as a function of the temperature and the concentration of oxygen and nitrogen while the change in the specific gravity was established as a function of the concentration of oxygen and nitrogen. Finally it was ascertained that in order to ensure long life for a 20 kW valve with a metallic anode it is necessary to have a strip of titanium of 14 sq. cm surface, weighing approx. 0.3 g, annealed at 800–800°C to fix the nitrogen, oxygen, carbon monoxide and carbon dioxide irreversibly, and a strip of titanium of 0.25 sq. cm surface, weighing approx. 0.007 g, annealed at a temperature not exceeding 400°C to fix the hydrogen irreversibly.

TOMASCHEK, Z.

G. FRANK, Mat nat Anz ung Akad Wiss, 1936, 54, 417-431

PROCEDURES AND PROPERTIES INDEX																									
1ST AND 2ND (2008)													1ST AND 2ND (2008)												
<p>Mechanism of the depolarization of local cathodes in corrosion involving oxygen depolarization. N. D. Tomashevsky. (Trans. 2nd Russ. Conf. Corrosion Metals, 1943, 2, 11-35).—The current strength between a Zn plate and a smaller Cu plate in 0.5N-HaCl, measured at various areas f of the Cu plate, is $I = k(2.25\sqrt{f} + f/d)$, k being a const. which can be calc. from the diffusion coeff. of O_2 and δ the thickness of the diffusion layer. Experiments give $\delta \sim 1$ mm. Only at $f > 0.5$ sq. cm. is $I \propto f$. If a Cu plate and a Zn ring surrounding it in the horizontal plane are covered with a layer, d cm. thick, of 3% NaCl, I has a sharp max. at $d \sim 0.08$ mm.; at lower d vals. I is smaller because of the increased ohmic resistance, and at higher d because of the longer diffusion path for O_2; at $d > 1$ mm. I is independent of d. I is raised by agitating the air above the NaCl solution. The term $2.25\sqrt{f}$ is due to the c.d. being larger near the edge of the cathode. Measurements of I between a Zn plate and a set of concentric Cu rings show that the c.d. 0-0.5 mm. from the edge is 4-5 times that in the centre of the cathode. If, at a const. f, the cathode consists of several spots instead of one plate, I in 3% NaCl is increased since the cross-section of the diffusion path for O_2 is raised, but the difference is negligible in N-HCl, in which H_2 is evolved. If the layer of 3% NaCl covering the cathode is < 1 mm. thick, the difference between a single and a composite cathode is reduced. If d is larger, an increase of the f of the composite cathode (or of the amount of cathodic inclusions in a corroding metal) does not cause a marked increase of I, the cross-section of the diffusion path being nearly equal to the total area of the corroding metal independently of the area of the cathodic inclusions.</p> <p style="text-align: right;">J. J. B.</p>																									
<p>ASH-324 METALLURGICAL LITERATURE CLASSIFICATION</p>																									

16.

Cathodic processes at an iron electrode under conditions of corrosion with oxygen depolarisation. N. D. Tomashev (Compt. rend. Acad. Sci. U.R.S.S., 1941, 28, 203-205). Cathodic polarisation curves on Fe surfaces supplied with O_2 in a bath at pH ~ 9 , show 3 sections corresponding with (1) $O_2 + 4e + 2H_2O \rightarrow 4OH^-$ up to c.d. ~ 1 mA. per sq. cm., (2) control of cathode reaction velocity by diffusion, up to -0.6 v., (3) discharge of H^+ above -0.6 v. An oxidised Fe cathode is less effective than a clean one. Fe_3O_4 is reduced when the potential is < -0.5 v. No H_2O_2 is formed. L. J. J.

30.14

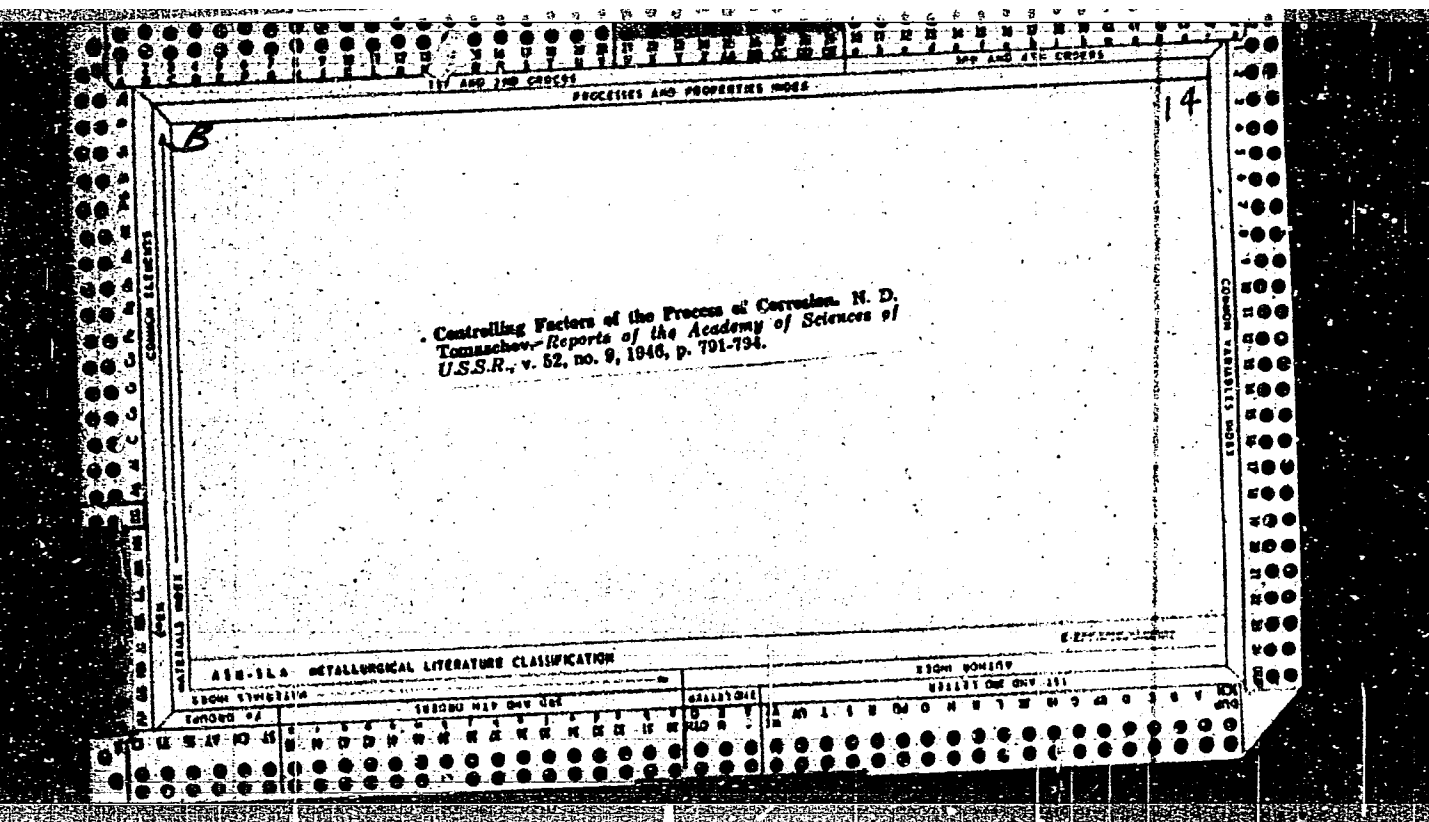
N. D. Tomashov

Graphical method of calculating polyelectrode electrochemical systems as applied to corrosion processes. N. D. Tomashov (*Compt. rend. Acad. Sci. U.R.S.S.*, 1941, 80, 621--623).—Comparison of cathode and anode polarisation curves makes it possible to calculate the work of each separate electrode in a polyelectrode system with any no. of electrodes. The relative surface area of each electrode in the system must be known, and the cathode and anode polarisation curves (c.d. against potential) under conditions close to the working conditions of the system must be plotted for each component of the system separately. The assumption is then made that the effective potentials of all the electrodes of the system become equal as the result of polarisation. This is usually the case if the total ohmic resistance is small. The c.d. potential curves are replotted on a general current-potential diagram. The general potential of the system is the potential at which the sum of all the cathodic currents equals the sum of all the anodic currents. The curve indicates which electrodes will act as cathodes and which as anodes. In some cases the polarisation curves for the individual electrodes can be calc. but they usually have to be obtained experimentally.

A. J. M.

COMMON ELEMENTS		PROPERTIES AND PROPERTIES INDEX		COMMON ELEMENTS	
<p>41</p> <p>*Theory of the Electrochemical Poly-Electrode Systems and Their Application to Corrosion Problems. I.—Potentials in Binary Systems. G. W. Akimov and N. D. Tomashev (Zhur. Fizich. Khimii (J. Phys. Chem.), 1938, 8, (6), 623-639).—[In Russian.] See abstract from a German source, Met. Abs., 1937, 4, 634.—S. G.</p>					
<p>ASB-SLA METALLURGICAL LITERATURE CLASSIFICATION</p>					
<p>1938-1940</p>		<p>1941-1945</p>		<p>1946-1950</p>	
<p>1951-1955</p>		<p>1956-1960</p>		<p>1961-1965</p>	
<p>1966-1970</p>		<p>1971-1975</p>		<p>1976-1980</p>	
<p>1981-1985</p>		<p>1986-1990</p>		<p>1991-1995</p>	
<p>1996-2000</p>		<p>2001-2005</p>		<p>2006-2010</p>	

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COMMON ELEMENTS

PROCESSES AND PROPERTIES INDEX

2-1

BC

Theory of many-electrode electrochemical systems and its application to corrosion problems. III. Many-electrode galvanic systems.
N. D. TOMASCHOV. (J. Phys. Chem. Russ., 1938, 12, 414—426).—The investigation previously described (A., 1938, I, 34) has been extended to include systems of 4 and 5 electrodes. In many-electrode systems inclusion of a new electrode as cathode increases the activity of all the anodes and diminishes that of the cathodes, whilst inclusion of a new anode increases the activity of the cathodes and diminishes that of the anodes. The difference effect (cf. A., 1936, 1474) and the action of protectors both seem to depend on diminution in the activity of local elements on the surface of the metal, as a result of diminished activity of the local cathodes in the first case, and lessened activity of the local anodes in the second. R. C.

OPEN

MATERIALS INDEX

ASTM-SLA METALLURGICAL LITERATURE CLASSIFICATION

AUTHOR

SUBJECT

RELATIONS

INDEXING

A	B	C	D	E	F	G	H	I	J	K	L	M	N	O	P	Q	R	S	T	U	V	W	X	Y	Z	AA	AB	AC	AD	AE	AF	AG	AH	AI	AJ	AK	AL	AM	AN	AO	AP	AQ	AR	AS	AT	AU	AV	AW	AX	AY	AZ	BA	BB	BC	BD	BE	BF	BG	BH	BI	BJ	BK	BL	BM	BN	BO	BP	BQ	BR	BS	BT	BU	BV	BW	BX	BY	BZ	CA	CB	CC	CD	CE	CF	CG	CH	CI	CJ	CK	CL	CM	CN	CO	CP	CQ	CR	CS	CT	CU	CV	CW	CX	CY	CZ	DA	DB	DC	DD	DE	DF	DG	DH	DI	DJ	DK	DL	DM	DN	DO	DP	DQ	DR	DS	DT	DU	DV	DW	DX	DY	DZ	EA	EB	EC	ED	EE	EF	EG	EH	EI	EJ	EK	EL	EM	EN	EO	EP	EQ	ER	ES	ET	EU	EV	EW	EX	EY	EZ	FA	FB	FC	FD	FE	FF	FG	FH	FI	FJ	FK	FL	FM	FN	FO	FP	FQ	FR	FS	FT	FU	FV	FW	FX	FY	FZ	GA	GB	GC	GD	GE	GF	GG	GH	GI	GJ	GK	GL	GM	GN	GO	GP	GQ	GR	GS	GT	GU	GV	GW	GX	GY	GZ	HA	HB	HC	HD	HE	HF	HG	HH	HI	HJ	HK	HL	HM	HN	HO	HP	HQ	HR	HS	HT	HU	HV	HW	HX	HY	HZ	IA	IB	IC	ID	IE	IF	IG	IH	II	IJ	IK	IL	IM	IN	IO	IP	IQ	IR	IS	IT	IU	IV	IW	IX
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COMMON ELEMENTS										PROCESSES AND PROPERTIES INDEX										SPECIFIC AND ATOM CATEGORIES									
<p>Be</p> <p>Theory of electrochemical systems with many electrodes and its application to corrosion problems. I. Potentials of binary systems. G. V. ARINOV and N. D. TOMASOV (J. Phys. Chem. Russ., 1936, 8, 622-639).—The potential of a "binary electrode," consisting, e.g., of Cu and Cd plates, depends on the relative situation and distances of the plates and the liquid junction, on the relative size of the plates, and on the polarization of the cathodic plate; a theory accounting for these effects is put forward. J. J. B.</p>																													
<p>ASM-A14 METALLURGICAL LITERATURE CLASSIFICATION</p>																													
<p>SECTION 1</p>										<p>SECTION 2</p>										<p>SECTION 3</p>									
<p>SECTION 1</p>										<p>SECTION 2</p>										<p>SECTION 3</p>									

1ST AND 2ND COLUMNS												3RD AND 4TH COLUMNS											
PROCESSING AND PROPERTY INDEX																							
<div style="position: relative;"> <div style="position: absolute; top: 10px; left: 10px; font-size: 2em; font-weight: bold;">B</div> <div style="position: absolute; top: 150px; left: 10px; transform: rotate(-90deg); font-size: 0.8em;">CROSS ELEMENTS</div> <div style="position: absolute; top: 650px; left: 10px; transform: rotate(-90deg); font-size: 0.8em;">MATERIALS INDEX</div> <div style="position: absolute; top: 350px; right: 10px; transform: rotate(90deg); font-size: 0.8em;">FROM LIBRARY NUMBER</div> <div style="position: absolute; top: 140px; left: 280px;"> <p>Change in the rate of dissolution of copper during anodic and cathodic polarizations.</p> <p>Difference effect. Protective effect. N. D. TOMASCHOV (Comp. rend. Acad. Sci. U.R.S.S., 1959, 24, 182-183).—The rate of dissolution of Cu in 0.1N and 0.1N aq. (NH₄)₂S₂O₈ and in 0.5N aq. CuCl₂ increases with increasing anodic polarization and decreases with increasing cathodic polarization. The rate of auto-dissolution of Cu (i.e., the rate of dissolution determined by the operation of local micro-elements) depends on the character of the current charge imposed, and for 0.1N aq. (NH₄)₂S₂O₈ it decreases with anodic polarization (positive difference effect) and cathodic polarization (positive protective effect). On the other hand, with 0.5N aq. CuCl₂ an increase in the rate of auto-dissolution of Cu is observed for both anodic (negative difference effect) and cathodic polarization (negative protective effect). The change of the potential of the Cu electrode during cathodic polarization is > that during anodic polarization. Dissolution of Cu (and probably other noble metals) in oxidizing media may thus be explained on the basis of the general electrochemical theory of the corrosion of metals.</p> <p style="text-align: right;">W. R. A.</p> </div> </div>																							
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PROCESSES AND PROPERTIES INDEX																			
<div style="position: absolute; top: 10%; left: 10%; font-size: 2em; font-weight: bold;">BC</div> <div style="position: absolute; top: 20%; right: 10%; font-size: 2em; font-weight: bold;">B I 6</div> <div style="position: absolute; top: 35%; left: 30%; font-size: 0.8em;"> <p>Mechanism of chromate film of anodic film on aluminum. N. D. ... (Soviet Acad. Sci. U.R.S.S., Ch. Sci. China, 1964) ... protection afforded by the ... with chromate ... decrease in porosity of the film due to hydration of Al_2O_3.</p> </div>																			
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COMMON ELEMENTS													PROCESSES AND PROPERTIES INDEX													COMMON ELEMENTS													PROCESSES AND PROPERTIES INDEX												
<p><i>BC</i></p> <p>Mechanism of corrosion of copper in electrolysis. N. D. TOMASCHOV (Compt. rend. Acad. Sci. U.R.S.S., 1939, 23, 649-653).—The Cu-Pt couple in 3% aq. NaCl acts only in presence of an oxidising agent (e.g., H_2O_2) at the cathode of the couple, Pt. Addition of H_2O_2 to the Pt compartment shifts the Pt potential in the positive direction to such an extent that dissolution of the Cu in the local element becomes possible. The dissolution of Cu in the couple $Cu 0.1N-(NH_4)_2SO_4 Pt$ occurs when $(NH_4)_2S_2O_8$ is added to the Pt compartment. Similar results have been obtained with the couple Cu-Pt in HNO_3 and H_2SO_4. These data indicate that the explanation of the mechanism of the dissolution of Cu (and other noble metals) based on a preliminary oxidation is not correct, and that the usual electrochemical theory of corrosion, after a detailed treatment of the processes of cathodic depolarisation, is in more complete agreement with the experimental data. W. R. A.</p>																										<p>ASB-SLA METALLURGICAL LITERATURE CLASSIFICATION</p>																									

BC

A I 1

Behaviour of a separate local cathode under conditions of oxygen depolarization. N. D. Tomashov (*Compt. rend. Acad. Sci. U.R.S.S.*, 1960, 27, 982-984).—On the assumption that the rate-controlling factor in the corrosion of a metal immersed in an aq. solution is the diffusion of O_2 from the surface of the metal (the actual corrosion reaction $2H + O + H_2O \rightarrow 2OH$ being very rapid), the following expression is deduced for the strength of the local current flowing as a result of the corrosion: $I = k\Delta C D \sqrt{(\pi/\delta)} \sin(\phi + \pi/4)$, where $k = \text{const.}$, ΔC is the difference of $[O_2]$ between the surface of the solution and that of the metal, D is the diffusion coeff. for O_2 , δ is the surface area of the local cathode, ϕ is the angle between the normal to the cathode surface and the generatrix of a cone of which a frustum is contained between the surfaces of solution and cathode (both assumed to be circular) which are at a vertical distance d apart. Theoretical and experimental val. of I are in good agreement for val. of d from 0.06 to 0.7 mm., the experimental figures showing negative deviations from the theoretical below 0.06 mm. since the ohmic resistance becomes the controlling factor for this film of solution, and positive ones above 0.7 mm. since convection as well as diffusional transfer of O_2 becomes important. From the theoretical expression it can be deduced that the c.d. at a local cathode will rise rapidly as the edge of the electrode is approached; the relation between the mean c.d. will thus not be directly \propto the area unless this is large but will be the greater the smaller is the area and the greater the perimeter/area ratio.
R. C. M.

ASH 514 METALLURGICAL LITERATURE CLASSIFICATION

3v ab

H. I. The Questions

Effect of distribution and density of local cathodes on the rate of corrosion under conditions of oxygen depolarization. N. D. Tomashov. (Compt. rend. Acad. Sci. U.S.S.R., 1940, 27, 991-994).
 —Since the mean e.d. across the cathodic part of a local couple under conditions of depolarization by O_2 diffusion increases as the area is reduced (preceding abstract), it is to be expected that the total current across a piece of corroding metal having cathodic inclusions will depend on their size and dispersion. If the total cathodic surface remains const. but its dispersion increases together with the area of the metal surface, the total current will increase because the area of electrolyte surface from which O_2 can diffuse to the cathodes will increase. A point is reached where the dispersion of the cathodic surface is so great that O_2 streams to the cathodic parts of the mosaic surface no longer overlap, and the total current then remains const. Experimental support for these conclusions has been obtained by measuring the change in total current between Cu(8% NaCl)/Zn couples when the operating area of the Cu was kept const. whilst the area of the mosaic in which it was located was decreased.
 R. C. M.

250. 260.

Subsidence of a separate local cathode under conditions of oxygen reduction. M. D. Tomassov (Compt. rend. Acad. Sci. U.R.S.S., 1940, 27, 983—986).—On the assumption that the rate-controlling factor in the corrosion of a metal immersed in an aq. solution is the diffusion of O_2 from the surface to the metal (the actual corrosion reaction $2H + O + H_2O \rightarrow 2OH$ being very rapid), the following expression is deduced for the strength of the local current flowing as a result of the corrosion: $I = kACD(\sqrt{v/d} - \tan \phi + f_0/d)$, where $k = \text{const.}$, ΔC is the difference of $[O_2]$ between the surface of the solution and that of the metal, D is the diffusion coeff. for O_2 , f_0 is the surface area of the local cathode, ϕ is the angle between the normal to the cathode surface and the generatrix of a cone of which a frustum is contained between the surfaces of solution and cathode (both assumed to be circular) which are at a vertical distance d apart. Theoretical and experimental vals. of I are in good agreement for vals. of d from 0.05 to 0.7 mm., the experimental figures showing negative deviations from the theoretical below 0.05 mm. since the ohmic resistance becomes the controlling factor for thin films of solution, and positive ones above 0.7 mm. since convectional as well as diffusional transfer of O_2 becomes important. From the theoretical expression it can be deduced that the c.d. at a local cathode will rise rapidly as the edge of the electrode is approached: the relation between the mean c.d. will thus not be directly \propto the area unless this is large but will be the greater the smaller is the area and the greater the perimeter/area ratio.
R. C. M.

1ST AND 2ND GROUPS PROCESSES AND PROPERTIES INDEX

BE

Effect of distribution and dispersity of local cathodes on the rate of corrosion under conditions of oxygen depolarization. N. D. Tomaschew. (Compt. rend. Acad. Sci. U.R.S.S., 1940, 87, 947-950).

—Since the mean c.d. across the cathodic part of a local couple under conditions of depolarization by O_2 diffusion increases as the area is reduced (preceding abstract), it is to be expected that the total current across a piece of corroding metal having cathodic inclusions will depend on their size and dispersion. If the total cathodic surface remains const. but its dispersion increases together with the area of the metal surface, the total current will increase because the area of electrolyte surface from which O_2 can diffuse to the cathodes will increase. A point is reached where the dis-

ASB-55A METALLURGICAL LITERATURE CLASSIFICATION

SECTION NUMBER

SECTION

SECTION

Rev. Abs.

W I - 26 Metals, Metallurgy, etc.

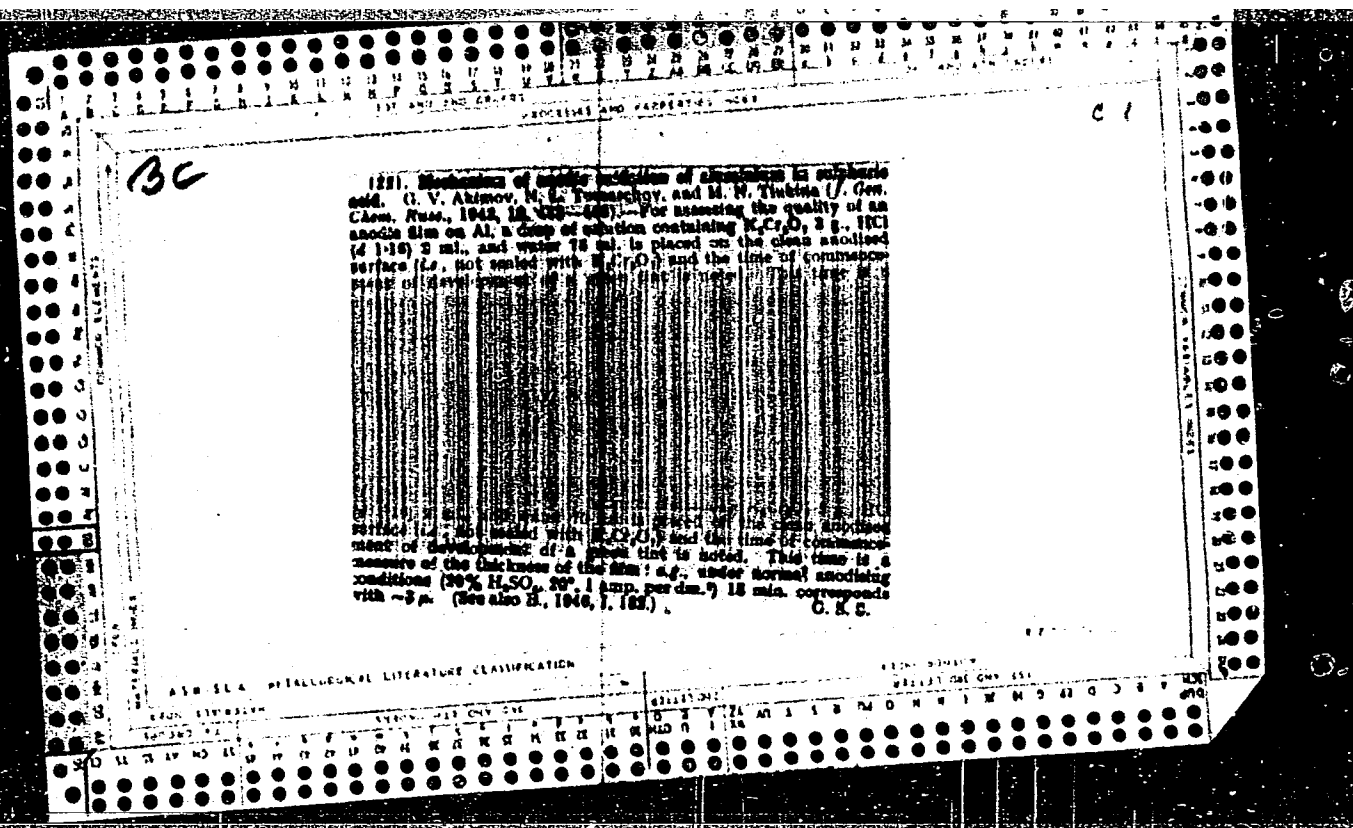
Cathodic processes in metallic corrosion. N. D. Tozashov
(*Compt. rend. Acad. Sci. U.R.S.S.*, 1966, 88, 181-184). An
analysis of theoretical curves of cathodic polarization in metallic
corrosion is presented. The location of crit. points on the curves
and the characteristic features at and between these points are
arranged in tabular form. C. R. H.

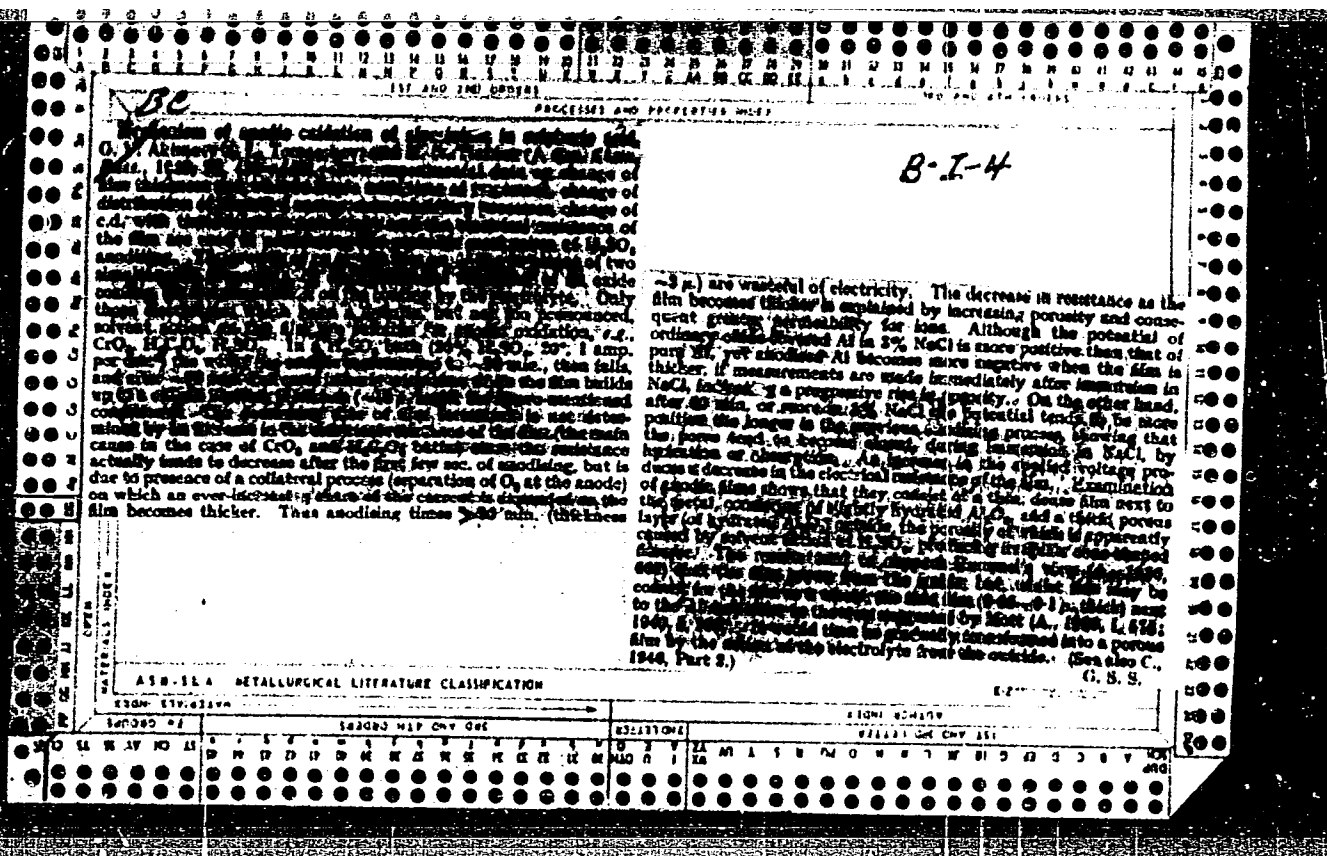
BC

A-1

Theory of many-electrode electrochemical systems and its application to corrosion problems. II. Three-electrode galvanic systems. N. D. TOMASCHOK. (J. Phys. Chem. Russ., 1937, 9, 43-83).—A Cd electrode was immersed in an electrolyte between a Pt anode and a Zn cathode and connected to various parts of the external circuit. Direction and strength of the current in the intermediate electrode were determined in relation to its position in the liquid, to the resistances of the external circuit on both sides of the branching, and to the conditions of electrolysis. E. R.

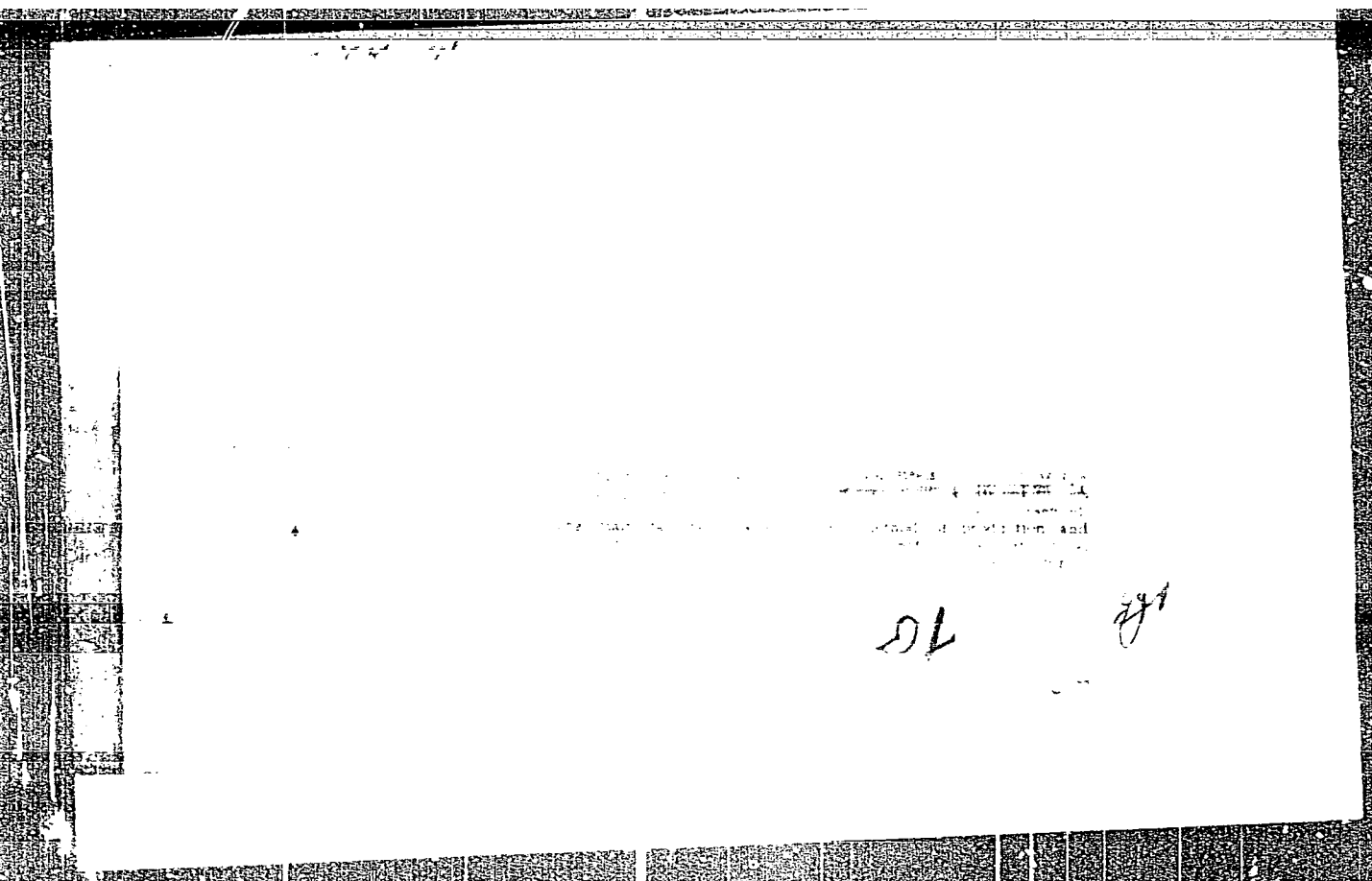
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CIA-RDP86-00513R001756210001-8



APPROVED FOR RELEASE: 04/03/2001

CIA-RDP86-00513R001756210001-8"

YUGOSLAVIA

Ivo TOMASEC [Affiliation not given.]

"The European Symposium on Diseases of Fish and Market Controls."

Belgrade, Veterinarski Glasnik, Vol 17, No 4, 1963; pp 386-389.

Abstract Report about this 5-day meeting held in Turin in Italy in October 1962. Resolutions about certain diseases of edible fish and about recommended inspection practices are given in considerable detail.

1/1

[illegible]

YUGOSLAVIA/Farm Animals - Honey Bee.

C-4

Abs Jour : Ref Zhur - Biol., No 1, 1959, 2767

Author : Tomanec, I.

Inst : -

Title : Effect of Certain Antibiotics on Bees.

Orig Pub : Veterin. arb., 1957, 27, No 3-4, 71-80.

Abstract : Terramycin (T), streptomycin (S) and penicillin (P) in sugar syrup (1:1) were fed to healthy bees taken from colonies and also to bees settled in groups of 20 to 100, each inside little cells at room temperature or at thermostatically controlled temperature of 30°. As compared with the control bees the greatest prolongation of life-span, improvement in the development of the brood, and increase of the honey yield to approximately double the normal, were obtained by applying S (0.25-0.5 grams per bee colony). The effect of P (200,000 units per colony)

Card 1/3

- 50 -

YUGOSLAVIA/Farm Animals - Honey Bee.

6-4

Abs Jour : Rev. Mirr - Biol., No 1, 1951, 2087

was less extensive, and inconsistent; in some families the development of brood and the construction of the honeycomb were improved. As for families fed with F for 20-30 days, observations showed merely an improvement in the development of the brood. In experiments with daged bees, the antibiotics had a weaker and less regular effect, although the bees given P did live somewhat longer than the control bees. The antibiotics exerted no effect on the development of salivary glands and the regeneration of the epithelial cells of the midgut. They modified appreciably the intestinal flora. Feeding with F was followed by the disappearance of the normal flora and by an intense development of yeast. Feeding with S was followed by the disappearance of minute bacteria, especially *Lact. curvica*, while feeding with P was followed by an intense proliferation of cocci and "pluton" forms.

Card 2/3

YUGOSLAVIA/Farm Animals - Honey Bee.

Q-4

Abs Jour : Rev Jour - Biol., No 1, 1956, 2767

It is probable that the modified flora affects metabolism and possibly also the enzyme system. -- V.A. Kanczyka

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 governing correct design are given. Soft or rein-
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